

Lewis Formulas (LF)

Shape

 e^- pairs on LF

Bonding units in LF

Rank polar bonds
choose most polar bondPeriodic Trends

a) size

b) ionization energy

Lab questions:

easy cation

easy anion

Flame tests

 K_2CrO_4 tests

Addition of 6M HCl

Rxn of NH_4OH with Fe^{+3} StoichiometryEF & MF by $\frac{MF_{Mm}}{EF_{Mm}} = \#(EF)$

EF by grams

EF by % mass

EF by moles

Mm calcs

mass $\rightarrow$ mole $\rightarrow$ moleculemolecule $\rightarrow$ mole $\rightarrow$ massTheoretical yield in
moles & mass

$$\% \text{ yield} = \frac{\text{act}}{\text{theo}} \times 100$$

Limiting Reactant

Mass other reactant

Mole ratios