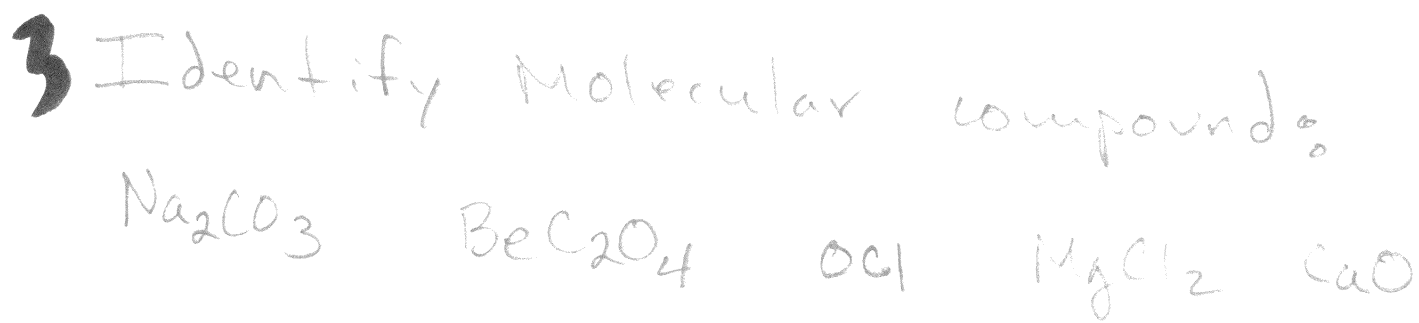
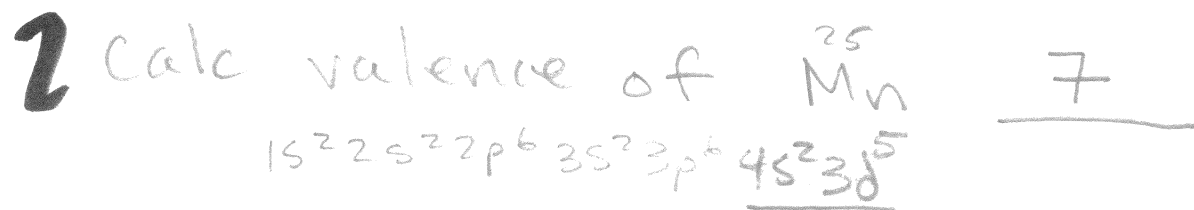
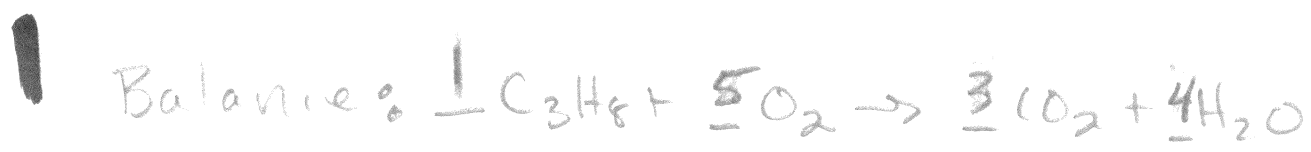


Draw the following
Formulas



Which pair of elements is
closest in size?

Cu, Ga, Se, Ge

Ga, Ge

Cu, Se

From which element is it
easier to remove an e^- ?

S

Te

8

Chem 7 white Board

(1)

~~100~~ ~~51~~ 9

Formula

barium sulfate — BaSO_4

hydrobromic acid — HBr

nitric acid — HNO_3

perchloric acid HClO_4

lead(II) acetate — $\text{Pb}(\text{C}_2\text{H}_3\text{O}_2)_2$

Silver nitrate — AgNO_3

copper (II) carbonate

→ CuCO_3

(2)

Name

~~so~~ ~~write Formula~~

1

NH_4Cl ammonium chloride

$\text{Cu}(\text{OH})_2$ copper (II) hydroxide

Fe_2S_3 iron(III) sulfide

$\text{Sn}_3(\text{PO}_4)_2$ tin (II) phosphate

AlPO_4 aluminum phosphate

~~1~~ ~~2~~ ~~3~~ ~~4~~ ~~5~~ ~~6~~ ~~7~~ ~~8~~ ~~9~~ 10

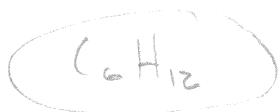
Which of the following atoms require $2e^-$ to fill outer shell?

~~11~~ ~~12~~ ~~13~~ ~~14~~ ~~15~~ ~~16~~ ~~17~~ ~~18~~ Ca, Se, Si, Cl, Ar

~~19~~ ~~20~~ ~~21~~ ~~22~~ ~~23~~ ~~24~~ ~~25~~ ~~26~~ ~~27~~ ~~28~~ ~~29~~ ~~30~~ ~~31~~ ~~32~~ ~~33~~ ~~34~~ ~~35~~ ~~36~~ ~~37~~ ~~38~~ ~~39~~ ~~40~~ ~~41~~ ~~42~~ ~~43~~ ~~44~~ ~~45~~ ~~46~~ ~~47~~ ~~48~~ ~~49~~ ~~50~~ ~~51~~ ~~52~~ ~~53~~ ~~54~~ ~~55~~ ~~56~~ ~~57~~ ~~58~~ ~~59~~ ~~60~~ ~~61~~ ~~62~~ ~~63~~ ~~64~~ ~~65~~ ~~66~~ ~~67~~ ~~68~~ ~~69~~ ~~70~~ ~~71~~ ~~72~~ ~~73~~ ~~74~~ ~~75~~ ~~76~~ ~~77~~ ~~78~~ ~~79~~ ~~80~~ ~~81~~ ~~82~~ ~~83~~ ~~84~~ ~~85~~ ~~86~~ ~~87~~ ~~88~~ ~~89~~ ~~90~~ ~~91~~ ~~92~~ ~~93~~ ~~94~~ ~~95~~ ~~96~~ ~~97~~ ~~98~~ ~~99~~ ~~100~~ ~~101~~ ~~102~~ ~~103~~ ~~104~~ ~~105~~ ~~106~~ ~~107~~ ~~108~~ ~~109~~ ~~110~~ ~~111~~ ~~112~~ ~~113~~ ~~114~~ ~~115~~ ~~116~~ ~~117~~ ~~118~~ ~~119~~ ~~120~~ ~~121~~ ~~122~~ ~~123~~ ~~124~~ ~~125~~ ~~126~~ ~~127~~ ~~128~~ ~~129~~ ~~130~~ ~~131~~ ~~132~~ ~~133~~ ~~134~~ ~~135~~ ~~136~~ ~~137~~ ~~138~~ ~~139~~ ~~140~~ ~~141~~ ~~142~~ ~~143~~ ~~144~~ ~~145~~ ~~146~~ ~~147~~ ~~148~~ ~~149~~ ~~150~~ ~~151~~ ~~152~~ ~~153~~ ~~154~~ ~~155~~ ~~156~~ ~~157~~ ~~158~~ ~~159~~ ~~160~~ ~~161~~ ~~162~~ ~~163~~ ~~164~~ ~~165~~ ~~166~~ ~~167~~ ~~168~~ ~~169~~ ~~170~~ ~~171~~ ~~172~~ ~~173~~ ~~174~~ ~~175~~ ~~176~~ ~~177~~ ~~178~~ ~~179~~ ~~180~~ ~~181~~ ~~182~~ ~~183~~ ~~184~~ ~~185~~ ~~186~~ ~~187~~ ~~188~~ ~~189~~ ~~190~~ ~~191~~ ~~192~~ ~~193~~ ~~194~~ ~~195~~ ~~196~~ ~~197~~ ~~198~~ ~~199~~ ~~200~~ ~~201~~ ~~202~~ ~~203~~ ~~204~~ ~~205~~ ~~206~~ ~~207~~ ~~208~~ ~~209~~ ~~210~~ ~~211~~ ~~212~~ ~~213~~ ~~214~~ ~~215~~ ~~216~~ ~~217~~ ~~218~~ ~~219~~ ~~220~~ ~~221~~ ~~222~~ ~~223~~ ~~224~~ ~~225~~ ~~226~~ ~~227~~ ~~228~~ ~~229~~ ~~230~~ ~~231~~ ~~232~~ ~~233~~ ~~234~~ ~~235~~ ~~236~~ ~~237~~ ~~238~~ ~~239~~ ~~240~~ ~~241~~ ~~242~~ ~~243~~ ~~244~~ ~~245~~ ~~246~~ ~~247~~ ~~248~~ ~~249~~ ~~250~~ ~~251~~ ~~252~~ ~~253~~ ~~254~~ ~~255~~ ~~256~~ ~~257~~ ~~258~~ ~~259~~ ~~260~~ ~~261~~ ~~262~~ ~~263~~ ~~264~~ ~~265~~ ~~266~~ ~~267~~ ~~268~~ ~~269~~ ~~270~~ ~~271~~ ~~272~~ ~~273~~ ~~274~~ ~~275~~ ~~276~~ ~~277~~ ~~278~~ 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~~102~~

A compound with $EFCH_2$ was found to have a Mm of 84. Find the chemical formula.



Q write the EC of outer valence of :

~~24~~



~~24~~



~~24~~



~~103~~

Q write EC of outer valence of :

~~24~~



~~24~~

~~24~~



~~24~~



12



~~13~~ ~~18~~ ~~47~~ ~~21~~ ~~13~~ ~~13~~

which groups of elements have the a radius most nearly the same? 13

- a. Be, B, C, N
- b. Ne, Ar, Kr, Xe
- c. Mg, Ca, Sr, Ba
- d. C, P, Se, I
- e. Cr, Mn, Fe, Co

~~14~~ ~~14~~ ~~14~~

write the correct EC for the following atoms or ions?

N^{-3}

$1s^2 2s^2 2p^6$

Na^+

$1s^2 2s^2 2p^6$

Mn

$1s^2 2s^2 2p^6 3s^2 3p^6 \underline{4s^2} 3d^5$
2 6 20

Mn^{+3}

$1s^2 2s^2 2p^6 3s^2 3p^6 3d^4$

14

~~24~~ Name the polyanions: ~~25~~

~~13~~

~~40~~



chlorate

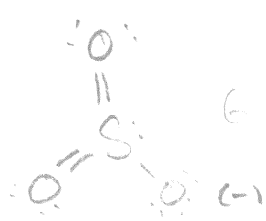
~~45~~

~~14~~



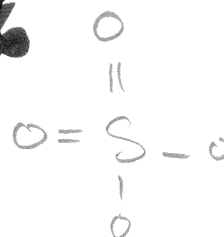
phosphate

~~15~~



SO_3^{2-}
 $(6 + 2(2)) + 2 = 26$
Sulfite

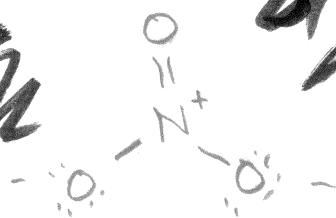
~~40~~



sulfate

~~12~~

~~5~~ ~~15~~



nitrate

~~12~~

~~15~~



carbonate

16

Draw LF

$5 + 6 + 1 = 12e^-$



17 Which compound is most ionic?

18 Select the compound with one unshared pair of e^- .

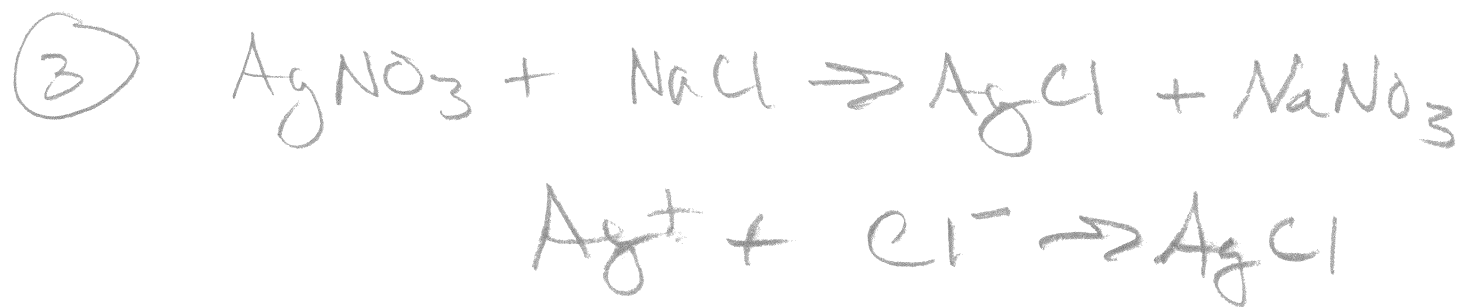
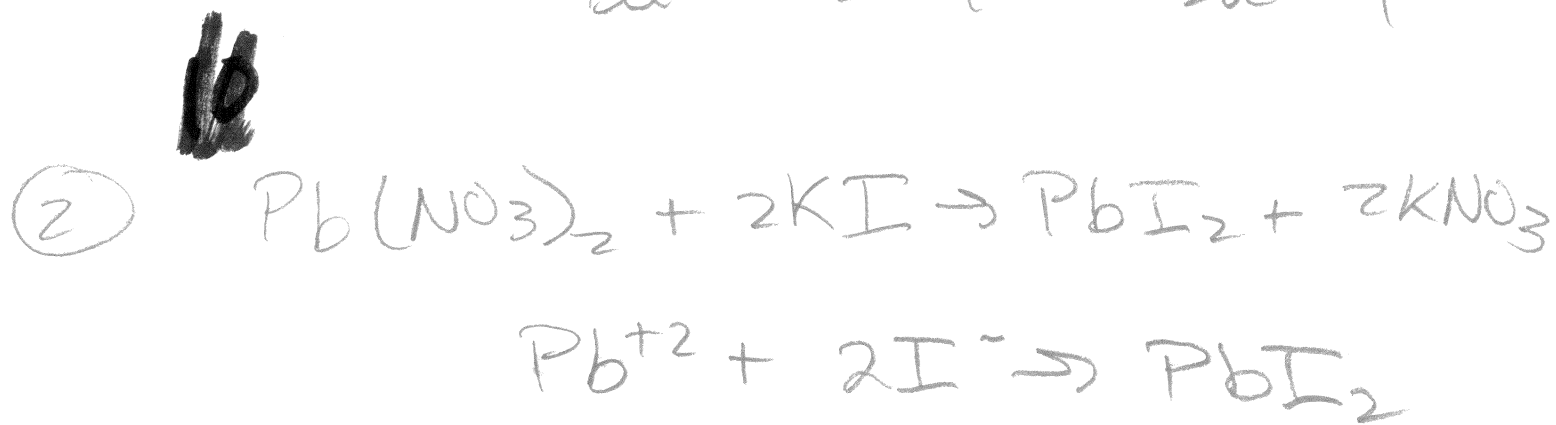
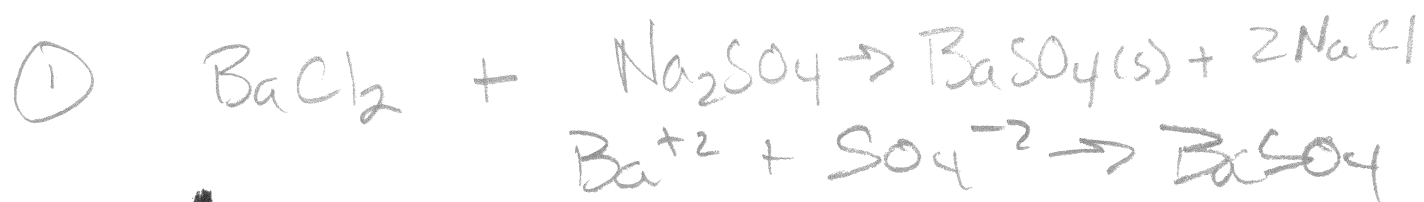
9/19

write the ionization equations

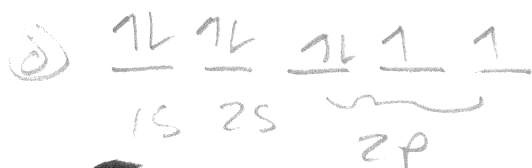
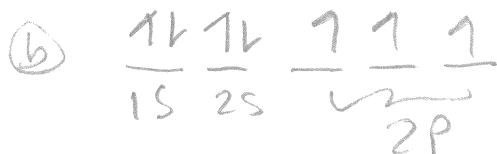
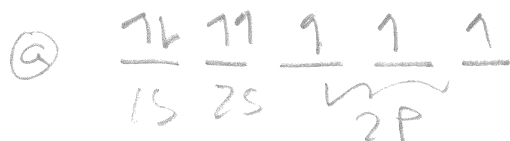


21 24 20

write the net ionic equation
for:



23) Indicate the OD that is allowed or not allowed by Pauli Exclusion Principle.



24) Se^{2-}

As

~~21~~ ~~21~~ 21

Use arrow + words to order Se, S, As in increasing at. radius.



22

25) Arrows + words to order following group in increasing IE.

Ar, Na, Cl, Al

Na Al Cl Ar

Small $\xrightarrow{\hspace{2cm}}$ big
low high

23 Calculate total e^- in O^{2-}
10

24 Calculate total e^- in Ba^{+2}
54

25 Write the outer valence EC of Sr
 $5s^2$

26 Calc the Z_{eff} of Sc's valence e^- .
 $Z_{eff} = 21 - 18 = +3$

27 Select molecule with 4 unshared pairs of e^- .



~~117~~ ~~119~~ ~~319~~
(28) Identify the element most likely to form a covalent bond with O. **28**

Na Cu Ti N Zn

~~117~~ ~~100~~ ~~40~~
(29) Identify the element that will form an ionic bond with O. **29**

P Se Br S K

~~118~~ ~~20~~ ~~121~~
(30) Identify the combustion equation.



~~112~~ ~~121~~ ~~412~~
(31) Write EC of valence of each ion:

~~21~~ Na^+ $2s^2 2p^6$

Cl^- $3s^2 3p^6$

S^{2-} $3s^2 3p^6$

Mg^{+2} $2s^2 2p^6$

30