

WB Chem 7 18-17-17

① Formula

Aluminum nitrate

Lead (IV) permanganate

Zinc chromate

Vanadium (II) phosphate

Silver hydroxide

② Arrange the following atoms in order of increasing reactivity using arrows and words:

Rb, Li, Na, K

③ Which atom is more reactive?

Ca or Na

④ The radius of the atom is determined by:
Zeff vs. core
Protons vs. Zeff
Zeff vs. Valence

⑤ Select OD that violates Hund's Rule: write the OD on your board!

$\frac{11}{15} \frac{11}{25} \frac{11}{25}$ $\frac{11}{15} \frac{11}{25} \frac{11}{25}$ $\frac{11}{15} \frac{11}{25} \frac{11}{25}$

⑥ Write the net ionic ppt rxn between:

AgNO_3 and NaCl

$\text{Pb(NO}_3)_2$ and KI

$\text{Ba(NO}_3)_2$ and Na_2SO_4

⑦ Write the ionization equations:

AgNO_3 , $\text{Pb(NO}_3)_2$, $\text{Ba(NO}_3)_2$

MgCl_2

⑧ When Rb reacts with oxygen, the charge on Rb is: _____

⑨ A soln is flame tested producing a deep red flame. The element is _____.

⑩ Choose the element with the lowest IE.

Cl, P, S, Si

⑪ An atom of Cl has 17e. How many energy levels does it need for its e^- ?

⑫ Name _____

Ca(OH)_2

$\text{Fe(NO}_3)_3$

$(\text{NH}_4)_2\text{C}_2\text{O}_7$

Cr(CN)_3

Li_2CO_3

⑬ Write the EC: He, Ne, Ar

⑭ Mg reacts with oxygen. Write the correct chemical formula for the compound that forms.

⑮ Name _____

Na_2CO_3

NaOH

MgBr_2

KCl

FeCl_2

⑯ Which element has the highest attractive forces?

Al, P, S, Si, Cl, Ar

⑰ Which element has the lowest attractive forces?

Cr, Mn, Sc, Ti, Fe, Zn

⑱ The smallest atom is: Bi, Sb, As, P, N

⑲ Select the atom with 2 single e^- in the highest occupied energy level:

Mg, O, N, F

⑳ Write the outer valence EC for F, Cl, Br.

㉑ Which of the following elements have 2 unpaired e^- in their EC?

㉒ Select the smallest element: Br, Se, As, Ge

㉓ Calc # core e^- in C

㉔ Calc # core e^- in Ar.

㉕ Name this poly anion: $\text{O}=\text{S}(\text{O})_2^-$

㉖ Write the ionization equation:

Al(OH)_3

SnCrO_4

HgBr_2

$\text{Cr}_2(\text{SO}_4)_3$

KMnO_4

Fe(OH)_3

㉗ Which compound is most ionic?

PCl_3 , OF_2 , BF_3 , LiF , SF_6

㉘ Given:

NaCl , MgBr_2 , LiOH , CuCO_3

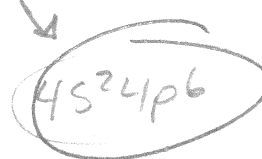
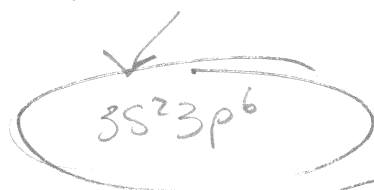
The dominant chemical bond is _____.

⑤ when calcium and bromine react, they form an ionic compound.

① write the symbol for each ion that forms: Ca^{+2} Br^-

② what is the chem. formula: CaCl_2

③ ECV for Ca^{+2} and Br^-



④ name: calcium bromide

2 Formula

14 aluminum nitrate $\text{Al}(\text{NO}_3)_3$

11 lead (IV) permanganate $\text{Pb}(\text{MnO}_4)_4$

zinc chromate ZnCrO_4

vanadium (II) phosphate $\text{V}_3(\text{PO}_4)_2$

silver hydroxide AgOH

copper (II) acetate $\text{Cu}(\text{C}_2\text{H}_3\text{O}_2)_2$

2 Arrange the following atoms in order of increasing reactivity.

Rb, Li, Na, K $Rb > K > Na > Li$

~~Li > Na > K > Rb~~

3 Which atom was more reactive?

Ca or Na

4 The radius of an atom is determined by which combination?

Z_{eff} vs core

#protons vs Z_{eff}

Z_{eff} vs valence

#protons vs core

5 Select the OD that violates Hund's rule:

$\frac{1\downarrow}{1s} \frac{1\downarrow}{2s} \frac{1\downarrow}{2p} \quad \underline{\hspace{1cm}} \quad \underline{\hspace{1cm}}$

$\frac{1\downarrow}{1s} \frac{1\downarrow}{2s} \frac{1\downarrow}{2p} \quad \underline{\hspace{1cm}} \quad \underline{\hspace{1cm}}$

$\frac{1\downarrow}{1s} \frac{1\downarrow}{2s} \frac{1\downarrow}{2p} \frac{1\downarrow}{2p} \frac{1\downarrow}{2p}$

$\frac{1\downarrow}{1s} \frac{1\downarrow}{2s} \frac{1\downarrow}{2p} \frac{1\downarrow}{2p} \frac{1\downarrow}{2p}$

6 write the net ionic ppt rxn for the rxn between:



7 write the ionization equation:



~~1~~ ~~2~~ ~~3~~ ~~4~~ ~~5~~ ~~6~~ ~~7~~ ~~8~~ ~~9~~ ~~10~~ ~~11~~ ~~12~~ ~~13~~ ~~14~~ ~~15~~ ~~16~~ ~~17~~ ~~18~~ ~~19~~ ~~20~~ ~~21~~ ~~22~~ ~~23~~ ~~24~~ ~~25~~ ~~26~~ ~~27~~ ~~28~~ ~~29~~ ~~30~~ ~~31~~ ~~32~~ ~~33~~ ~~34~~ ~~35~~ ~~36~~ ~~37~~ ~~38~~ ~~39~~ ~~40~~ ~~41~~ ~~42~~ ~~43~~ ~~44~~ ~~45~~ ~~46~~ ~~47~~ ~~48~~ ~~49~~ ~~50~~ ~~51~~ ~~52~~ ~~53~~ ~~54~~ ~~55~~ ~~56~~ ~~57~~ ~~58~~ ~~59~~ ~~60~~ ~~61~~ ~~62~~ ~~63~~ ~~64~~ ~~65~~ ~~66~~ ~~67~~ ~~68~~ ~~69~~ ~~70~~ ~~71~~ ~~72~~ ~~73~~ ~~74~~ ~~75~~ ~~76~~ ~~77~~ ~~78~~ ~~79~~ ~~80~~ ~~81~~ ~~82~~ ~~83~~ ~~84~~ ~~85~~ ~~86~~ ~~87~~ ~~88~~ ~~89~~ ~~90~~ ~~91~~ ~~92~~ ~~93~~ ~~94~~ ~~95~~ ~~96~~ ~~97~~ ~~98~~ ~~99~~ ~~100~~ ~~101~~ ~~102~~ ~~103~~ ~~104~~ ~~105~~ ~~106~~ ~~107~~ ~~108~~ ~~109~~ ~~110~~ ~~111~~ ~~112~~ ~~113~~ ~~114~~ ~~115~~ ~~116~~ ~~117~~ ~~118~~ ~~119~~ ~~120~~ ~~121~~ ~~122~~ ~~123~~ ~~124~~ ~~125~~ ~~126~~ ~~127~~ ~~128~~ ~~129~~ ~~130~~ ~~131~~ ~~132~~ ~~133~~ ~~134~~ ~~135~~ ~~136~~ ~~137~~ ~~138~~ 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~~1345~~ ~~1346~~ ~~134~~

Directions: Name the compound. Then write the ionization equation.

1. AuBr_3 gold(III) bromide

2. MgHPO_4 magnes. hydrophosphate

3. Ag_2SO_4 silver sulfate

4. $\text{Ca}(\text{OH})_2$ calcium hydroxide

5. $\text{Fe}(\text{NO}_2)_3$ iron(III) nitrite

6. $(\text{NH}_4)_2\text{Cr}_2\text{O}_7$ ammonium dichromate

7. $\text{Cr}(\text{CN})_3$ chromium(III) cyanide

8. Li_2CO_3 lithium carbonate

9. BaSO_4 barium sulfate

10. $\text{Mn}(\text{OH})_2$ manganese(II) hydroxide

13

Q1 Write the EC of He, Ne, Ar.

~~15~~ He $1s^2$

Ne $1s^2 2s^2 2p^6$

~~16~~ Ar $1s^2 2s^2 2p^6 3s^2 3p^6$

Q2 Principle quantum number (n) refers to shape of the orbital.

Q3 Principle quantum number (m_l) refers to orientation of the orbital.

Q4 ~~18~~ neutral ~~15~~ An atom with at. number equal to 37 has 37 electrons.

+1 charge $36 e^-$

~~16~~ Q5 When Mg reacts with nitrogen, Mg ionizes to form a charge of +2.



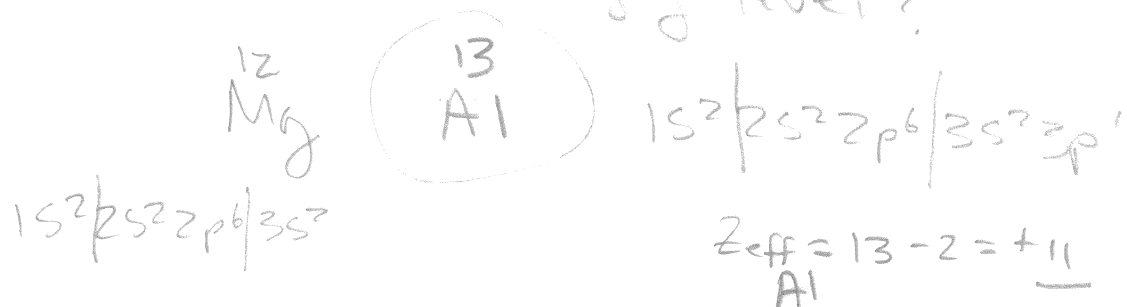
Naming Ionic Compounds Worksheet One

Give the name of the following ionic compounds:

- 1) Na_2CO_3 sodium carbonate
- 2) NaOH sodium hydroxide
- 3) MgBr_2 magnesium bromide
- 4) KCl pot chloride
- 5) FeCl_2 iron (II) chloride
- 6) FeCl_3 iron (III) chloride
- 7) Zn(OH)_2 zinc hydroxide
- 8) BeSO_4 beryllium sulfate
- 9) CrF_2 ~~cr~~ chromium (II) fluoride
- 10) Al_2S_3 Alu sulfide
-
- 11) PbO lead (II) oxide
- 12) Li_3PO_4 lithium phosphate
- 13) TiI_4 titanium (IV) iodide
- 14) Co_3N_2 cobalt (I) nitride
- 15) Mg_3P_2 mag phosphide
- 16) $\text{Ga(NO}_2)_3$ gallium nitrite
- 17) Ag_2SO_3 silver sulfite
- 18) NH_4OH ammonium hydroxide
- 19) Al(CN)_3 alu cyanide
- 20) $\text{Be(CH}_3\text{COO)}_2$ bery. acetate
 $\text{C}_2\text{H}_3\text{O}_2$

12 ~~13~~ ~~14~~ ~~15~~ ~~16~~ ~~17~~ ~~18~~ ~~19~~ ~~20~~ ~~21~~ ~~22~~ ~~23~~ ~~24~~ ~~25~~ ~~26~~ ~~27~~ ~~28~~ ~~29~~ ~~30~~ ~~31~~ ~~32~~ ~~33~~ ~~34~~ ~~35~~ ~~36~~ ~~37~~ ~~38~~ ~~39~~ ~~40~~ ~~41~~ ~~42~~ ~~43~~ ~~44~~ ~~45~~ ~~46~~ ~~47~~ ~~48~~ ~~49~~ ~~50~~ ~~51~~ ~~52~~ ~~53~~ ~~54~~ ~~55~~ ~~56~~ ~~57~~ ~~58~~ ~~59~~ ~~60~~ ~~61~~ ~~62~~ ~~63~~ ~~64~~ ~~65~~ ~~66~~ ~~67~~ ~~68~~ ~~69~~ ~~70~~ ~~71~~ ~~72~~ ~~73~~ ~~74~~ ~~75~~ ~~76~~ ~~77~~ ~~78~~ ~~79~~ ~~80~~ ~~81~~ ~~82~~ ~~83~~ ~~84~~ ~~85~~ ~~86~~ ~~87~~ ~~88~~ ~~89~~ ~~90~~ ~~91~~ ~~92~~ ~~93~~ ~~94~~ ~~95~~ ~~96~~ ~~97~~ ~~98~~ ~~99~~ ~~100~~ ~~101~~ ~~102~~ ~~103~~ ~~104~~ ~~105~~ ~~106~~ ~~107~~ ~~108~~ ~~109~~ ~~110~~ ~~111~~ ~~112~~ ~~113~~ ~~114~~ ~~115~~ ~~116~~ ~~117~~ ~~118~~ ~~119~~ ~~120~~ ~~121~~ ~~122~~ ~~123~~ ~~124~~ ~~125~~ ~~126~~ ~~127~~ ~~128~~ ~~129~~ ~~130~~ ~~131~~ ~~132~~ ~~133~~ ~~134~~ ~~135~~ ~~136~~ ~~137~~ ~~138~~ ~~139~~ ~~140~~ ~~141~~ ~~142~~ ~~143~~ ~~144~~ ~~145~~ ~~146~~ ~~147~~ 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~~773~~ ~~774~~ ~~775~~ ~~776~~ ~~777~~ ~~778~~ ~~779~~ ~~780~~ ~~781~~ ~~782~~ ~~783~~ ~~784~~ ~~785~~ ~~786~~ ~~787~~ ~~788~~ ~~789~~ ~~790~~ ~~791~~ ~~792~~ ~~793~~ ~~794~~ ~~795~~ ~~796~~ ~~797~~ ~~798~~ ~~799~~ ~~800~~ ~~801~~ ~~802~~ ~~803~~ ~~804~~ ~~805~~ ~~806~~ ~~807~~ ~~808~~ ~~809~~ ~~810~~ ~~811~~ ~~812~~ ~~813~~ ~~814~~ ~~815~~ ~~816~~ ~~817~~ ~~818~~ ~~819~~ ~~820~~ ~~821~~ ~~822~~ ~~823~~ ~~824~~ ~~825~~ ~~826~~ ~~827~~ ~~828~~ ~~829~~ ~~830~~ ~~831~~ ~~832~~ ~~833~~ ~~834~~ ~~835~~ ~~836~~ ~~837~~ ~~838~~ ~~839~~ ~~840~~ ~~841~~ ~~842~~ ~~843~~ ~~844~~ ~~845~~ ~~846~~ ~~847~~ ~~848~~ ~~849~~ ~~850~~ ~~851~~ ~~852~~ ~~853~~ ~~854~~ ~~855~~ ~~856~~ ~~857~~ ~~858~~ ~~859~~ ~~860~~ ~~861~~ ~~862~~ ~~863~~ ~~864~~ ~~865~~ ~~866~~ ~~867~~ ~~868~~ ~~869~~ ~~870~~ ~~871~~ ~~872~~ ~~873~~ ~~874~~ ~~875~~ ~~876~~ ~~877~~ ~~878~~ ~~879~~ ~~880~~ ~~881~~ ~~882~~ ~~883~~ ~~884~~ ~~885~~ ~~886~~ ~~887~~ ~~888~~ ~~889~~ ~~890~~ ~~891~~ ~~892~~ ~~893~~ ~~894~~ ~~895~~ ~~896~~ ~~897~~ 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~~1020~~ ~~1021~~ ~~1022~~ ~~1023~~ ~~1024~~ ~~1025~~ ~~1026~~ ~~1027~~ ~~1028~~ ~~1029~~ ~~1030~~ ~~1031~~ ~~1032~~ ~~1033~~ ~~1034~~ ~~1035~~ ~~1036~~ ~~1037~~ ~~1038~~ ~~1039~~ ~~1040~~ ~~1041~~ ~~1042~~ ~~1043~~ ~~1044~~ ~~1045~~ ~~1046~~ ~~1047~~ ~~1048~~ ~~1049~~ ~~1050~~ ~~1051~~ ~~1052~~ ~~1053~~ ~~1054~~ ~~1055~~ ~~1056~~ ~~1057~~ ~~1058~~ ~~1059~~ ~~1060~~ ~~1061~~ ~~1062~~ ~~1063~~ ~~1064~~ ~~1065~~ ~~1066~~ ~~1067~~ ~~1068~~ ~~1069~~ ~~1070~~ ~~1071~~ ~~1072~~ ~~1073~~ ~~1074~~ ~~1075~~ ~~1076~~ ~~1077~~ ~~1078~~ ~~1079~~ ~~1080~~ ~~1081~~ ~~1082~~ ~~1083~~ ~~1084~~ ~~1085~~ ~~1086~~ ~~1087~~ ~~1088~~ ~~1089~~ ~~1090~~ ~~1091~~ ~~1092~~ ~~1093~~ ~~1094~~ ~~1095~~ ~~1096~~ ~~1097~~ ~~1098~~ ~~1099~~ ~~1100~~ ~~1101~~ ~~1102~~ ~~1103~~ ~~1104~~ ~~1105~~ ~~1106~~ ~~1107~~ ~~1108~~ ~~1109~~ ~~1110~~ ~~1111~~ ~~1112~~ ~~1113~~ ~~1114~~ ~~1115~~ ~~1116~~ ~~1117~~ ~~1118~~ ~~1119~~ ~~1120~~ ~~1121~~ ~~1122~~ ~~1123~~ ~~1124~~ ~~1125~~ ~~1126~~ ~~1127~~ ~~1128~~ ~~1129~~ ~~1130~~ 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~~1353~~ ~~1354~~ ~~1355~~

13 which element, Mg or Al, requires more ionization energy to remove an e^- from the $n=2$ energy level?



$$Z_{\text{eff Mg}} = 12 - 2 = \underline{+10}$$

why: Al has higher
nuc. charge

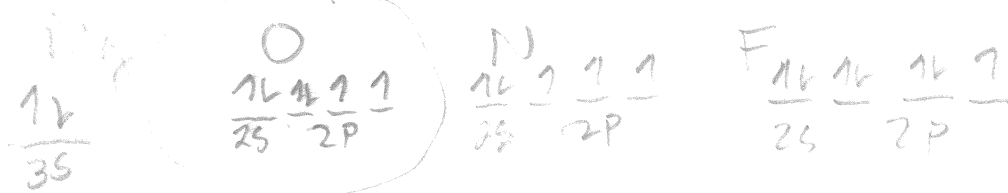
Calculate the Z_{eff} of the $n=2$ electrons of Mg and Al

$$Mg = +10$$

$$A1 = +11$$

The gas is hydrogen.

Select the atom with 2 single electrons in the highest occupied energy level.



20

Q1 Write EC for F, Cl, Br
and underline outer valence



21

21

21

Q2 Which of the following elements
have 2 unpaired e^- ?

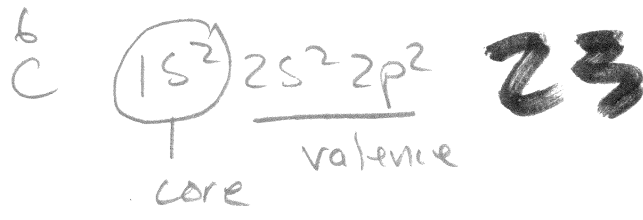


Select the smallest element from

Br, Se, As, Ge

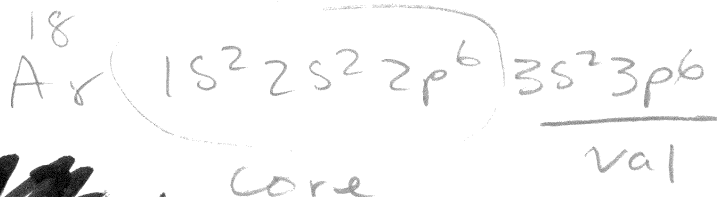
Given C with atomic number 6,

Calculate the core electrons.



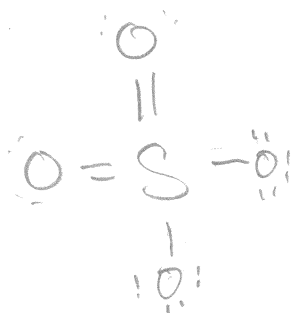
(2)

Given Ar, calculate the core e-



(10)

Name this polyanion:



Sulfate

Ionize

~~26~~



~~27~~

~~27~~

~~27~~

which compound is most ionic?



~~28~~

~~28~~

~~28~~

The type of bond in the following compounds is a _____ bond.



covalent

Ionic